NAME_____

WORKSHOP 9A: Intermolecular Forces Section _____

For the first part of this workshop, identify the type of crystal structure (Ionic, Molecular Polar, Molecular Nonpolar, Network-Covalent, or Metallic) present. Then determine the type of binding forces present in each (Ionic Bonds, Covalent Bonds, Metallic Bonds, London Dispersion Forces, Dipole Forces, and/or Hydrogen Bonds).

Substance	Type of Crystal	Type of Binding Force(s)
Ar		
CH_3Cl		
CH_3OH		
BCl_3		
CH_3OCH_3		
HF		
Hg		
N_2		
SiC		
CH_3COOH		
Diamond		

Circle the species with the higher boiling point and *briefly* justify your choice below.

1) Kr_ Justification: or Xe _____

2) C_2H_5OH_____
Justification: or CH_3OCH_3 _____

3) NaF_____
Justification: or MgO _____

4) N_2 _____
Justification: or NO _____

5) CH_4 _____
Justification: or SiH_4 _____

6) HF_ Justification:
or HI _____

7) CO_2 _____
Justification:
or NH_3 _____

8) CH_4 _____
Justification:
or CCl_4 _____

9) Cr_ Justification:
or Si _____

10) H_2O_____
Justification:
or SiO_2 _____

11) MgO_____
Justification:
or BaO _____

12) $CH_3CH_2CH_2CH_2CH_3$ _____
Justification:
or $(CH_3)_2CHCH_2CH_2CH_3$ _____