Standard Reduction Potentials in Aqueous Solution at 25°C	
Acidic Solution	**E° (V)**
$F_2(g) + 2\ e^- \longrightarrow 2\ F^-(aq)$	+2.87
$Co^{3+}(aq) + e^- \longrightarrow Co^{2+}(aq)$	+1.82
$Pb^{4+}(aq) + 2\ e^- \longrightarrow Pb^{2+}(aq)$	+1.8
$H_2O_2(aq) + 2\ H^+(aq) + 2\ e^- \longrightarrow 2\ H_2O$	+1.77
$NiO_2(s) + 4\ H^+(aq) + 2\ e^- \longrightarrow Ni^{2+}(aq) + 2\ H_2O$	+1.7
$PbO_2(s) + SO_4^{2-}(aq) + 4\ H^+(aq) + 2\ e^- \longrightarrow PbSO_4(s) + 2\ H_2O$	+1.685
$Au^+(aq) + e^- \longrightarrow Au(s)$	+1.68
$2\ HClO(aq) + 2\ H^+(aq) + 2\ e^- \longrightarrow Cl_2(g) + 2\ H_2O$	+1.63
$MnO_4^-(aq) + 8\ H^+(aq) + 5\ e^- \longrightarrow Mn^{2+}(aq) + 4\ H_2O$	+1.51
$Au^{3+}(aq) + 3\ e^- \longrightarrow Au(s)$	+1.5
$ClO_3^-(aq) + 6\ H^+(aq) + 5\ e^- \longrightarrow 1/2\ Cl_2(g) + 3\ H_2O$	+1.47
$BrO_3^- + 6\ H^+(aq) + 6\ e^- \longrightarrow Br^-(aq) + 3\ H_2O$	+1.44
$Cl_2(g) + 2\ e^- \longrightarrow 2\ Cl^-(aq)$	+1.358
$Cr_2O_7^{2-} + 14\ H^+(aq) + 6\ e^- \longrightarrow 2\ Cr^{3+}(aq) + 7\ H_2O$	+1.33
$MnO_2(s) + 4\ H^+(aq) + 2\ e^- \longrightarrow Mn^{2+}(aq) + 2\ H_2O$	+1.23
$O_2(g) + 4\ H^+(aq) + 4\ e^- \longrightarrow 2\ H_2O$	+1.229
$IO_3^-(aq) + 6\ H^+(aq) + 5\ e^- \longrightarrow 1/2\ I_2(aq) + 3\ H_2O$	+1.195
$ClO_4^-(aq) + 2\ H^+(aq) + 2\ e^- \longrightarrow ClO_3^-(aq) + H_2O$	+1.19
$Br_2(l) + 2\ e^- \longrightarrow 2\ Br^-(aq)$	+1.066
$AuCl_4^- + 3\ e^- \longrightarrow Au(s) + 4\ Cl^-(aq)$	+1
$NO_3^-(aq) + 4\ H^+(aq) + 3\ e^- \longrightarrow NO(g) + 2\ H_2O$	+0.96
$NO_3^-(aq) + 3\ H^+(aq) + 2\ e^- \longrightarrow HNO_2(aq) + H_2O$	+0.94
$2\ Hg^{2+}(aq) + 2\ e^- \longrightarrow Hg_2^{2+}(aq)$	+0.92
$Hg^{2+}(aq) + 2\ e^- \longrightarrow Hg(l)$	+0.855
$Ag^+(aq) + e^- \longrightarrow Ag(s)$	+0.7994
$Hg_2^{2+}(aq) + 2\ e^- \longrightarrow 2\ Hg(l)$	+0.789
$Fe^{3+}(aq) + e^- \longrightarrow Fe^{2+}(aq)$	+0.771
$SbCl_6^-(aq) + 2\ e^- \longrightarrow SbCl_4^-(aq) + 2\ Cl^-(aq)$	+0.75
$[PtCl_4]^{2-}(aq) + 2\ e^- \longrightarrow Pt(s) + 4\ Cl^-(aq)$	+0.73
$O_2(g) + 2\ H^+(aq) + 2\ e^- \longrightarrow H_2O_2(aq)$	+0.682
$[PtCl_6]^{2-}(aq) + 2\ e^- \longrightarrow [PtCl_4]^{2-}(aq) + 2\ Cl^-(aq)$	+0.68
$H_3AsO_4(aq) + 2\ H^+(aq) + 2\ e^- \longrightarrow H_3AsO_3(aq) + H_2O$	+0.58
$I_2(s) + 2\ e^- \longrightarrow 2\ I^-(aq)$	+0.535

$TeO_2(s) + 4\ H^+(aq) + 4\ e^- \longrightarrow Te(s) + 2\ H_2O$	+0.529
$Cu^+(aq) + e^- \longrightarrow Cu(s)$	+0.521
$Cu^{2+}(aq) + 2\ e^- \longrightarrow Cu(s)$	+0.337
$HgCl_2(s) + 2\ e^- \longrightarrow 2\ Hg(l) + 2\ Cl^-(aq)$	+0.27
$AgCl(s) + e^- \longrightarrow Ag(s) + Cl^-(aq)$	+0.222
$SO_4^{2-}(aq) + 4\ H^+(aq) + 2\ e^- \longrightarrow SO_2(g) + 2\ H_2O$	+0.2
$SO_4^{2-}(aq) + 4\ H^+(aq) + 2\ e^- \longrightarrow H_2SO_3(g) + H_2O$	+0.17
$Cu^{2+}(aq) + e^- \longrightarrow Cu^+(aq)$	+0.153
$Sn^{4+}(aq) + 2\ e^- \longrightarrow Sn^{2+}(aq)$	+0.15
$S(s) + 2\ H^+(aq) + 2\ e^- \longrightarrow H_2S(aq)$	+0.14
$AgBr(s) + e^- \longrightarrow Ag(s) + Br^-(aq)$	+0.0713
$2\ H^+(aq) + 2\ e^- \longrightarrow H_2(g)$ (reference electrode)	0
$Pb^{2+}(aq) + 2\ e^- \longrightarrow Pb(s)$	-0.126
$Sn^{2+}(aq) + 2\ e^- \longrightarrow Sn(s)$	-0.14
$AgI(s) + e^- \longrightarrow Ag(s) + I^-(aq)$	-0.15
$Ni^{2+}(aq) + 2\ e^- \longrightarrow Ni(s)$	-0.25
$Co^{2+}(aq) + 2\ e^- \longrightarrow Co(s)$	-0.28
$Tl^+(aq) + e^- \longrightarrow Tl(s)$	-0.34
$PbSO_4(s) + 2\ e^- \longrightarrow Pb(s) + SO_4^{2-}(aq)$	-0.356
$Se(s) + 2\ H^+(aq) + 2\ e^- \longrightarrow H_2Se(aq)$	-0.4
$Cd^{2+}(aq) + 2\ e^- \longrightarrow Cd(s)$	-0.403
$Cr^{3+}(aq) + e^- \longrightarrow Cr^{2+}(aq)$	-0.41
$Fe^{2+}(aq) + 2\ e^- \longrightarrow Fe(s)$	-0.44
$Ga^{3+}(aq) + 3\ e^- \longrightarrow Ga(s)$	-0.53
$Cr^{3+}(aq) + 3\ e^- \longrightarrow Cr(s)$	-0.74
$Zn^{2+}(aq) + 2\ e^- \longrightarrow Zn(s)$	-0.763
$2H_2O(l) + 2\ e^- \longrightarrow H_2(g) + 2OH^-(aq)$	-0.8277
$Cr^{2+}(aq) + 2\ e^- \longrightarrow Cr(s)$	-0.91
$Mn^{2+}(aq) + 2\ e^- \longrightarrow Mn(s)$	-1.18
$V^{2+}(aq) + 2\ e^- \longrightarrow V(s)$	-1.18
$Zr^{4+}(aq) + 4\ e^- \longrightarrow Zr(s)$	-1.53
$Al^{3+}(aq) + 3\ e^- \longrightarrow Al(s)$	-1.66
$H_2(g) + 2\ e^- \longrightarrow 2\ H^-(aq)$	-2.25
$Mg^{2+}(aq) + 2\ e^- \longrightarrow Mg(s)$	-2.37
$Na^+(aq) + e^- \longrightarrow Na(s)$	-2.714