Solubility Chart:

Combination of cations with anions either produce no precipitate, (i.e. the ions do not combine in solution) or they form a precipitate, leaving less than 1% of the ions in solution. Soluble combinations are labeled **(aq)**; insoluble combinations are labeled **(s).** Some combinations form gases or undergo complex reactions, these are labeled d, indicating decomposition. Spaces labeled ---- indicate no reference was found.

	$C_2H_3O_2^-$	AsO_4^{3-}	Br^-	CO_3^{2-}	Cl^-	CrO_4^{2-}	OH^-	I^-	NO_3^-	$C_2O_4^{2-}$	O^{2-}	PO_4^{3-}	SO_4^{2-}	S^{2-}	SO_3^{2-}
Group I & NH_4^+	(aq)	(aq)	(aq)	(aq)	(aq)	(aq)	(aq)	(aq)	(aq)	(aq)	(aq)	(aq)	(aq)	(aq)	(aq)
Al^{3+}	(aq)	(s)	(aq)	------	(aq)	------	(s)	(aq)	(aq)	------	(s)	(s)	(aq)	d	------
Ba^{2+}	(aq)	(s)	(aq)	(s)	(aq)	(s)	(aq)	(aq)	(aq)	(s)	(aq)	(s)	(s)	d	(s)
Bi^{3+}	--------	(aq)	d	(s)	d	-------	(s)	(s)	d	(s)	s)	(aq)	d	(s)	------
Ca^{2+}	(aq)	(s)	(aq)	(s)	(aq)	(aq)	(s)	(aq)	(aq)	(s)	(s)	(s)	(s)	d	(s)
Co^{2+} Ni^{2+} and Cu^{2+}	(aq)	(s)	(aq)	(s)	(aq)	(s)	(s)	(aq)	(aq)	(s)	(s)	(s)	(aq)	(s)	(s)
Fe^{2+}	(aq)	(s)	(aq)	(s)	(aq)	------	(s)	(aq)	(aq)	(s)	(s)	(s)	(aq)	(s)	(s)
Fe^{3+}	(s)	(s)	(aq)	-----	(aq)	------	(s)	------	(aq)	(aq)	(s)	(s)	(aq)	(s)	-------
Pb^{2+}	(aq)	(s)	(s)	(s)	(s)	(s)	(s)	(s)	(aq)	(s)	(s)	(s)	(s)	(s)	(s)
Mg^{2+}	(aq)	d	(aq)	(s)	(aq)	(aq)	(s)	(aq)	(aq)	(s)	(s)	(s)	(aq)	d	(aq)
Hg^{2+}	(aq)	(s)	(s)	(s)	(aq)	(aq)	(s)	(s)	(aq)	(s)	(s)	(s)	d	(s)	-----
Ag^+	(aq)	(s)	(s)	(s)	(s)	(s)	(s) -	(s)	(aq)	(s)	(s)	(s)	(s)	(s)	(s)
Zn^{2+}	(aq)	(s)	(aq)	(s)	(aq)	(s)	(s)	(aq)	(aq)	(s)	(s)	(s)	(aq)	(s)	(s)

It is useful to remember that all compounds containing Group I A ions are soluble. Also, all compounds containing the ammonium ion or the nitrate ion are soluble.